Chemistry Study Materials for Class 9

(MCQ Type Questions – Answers)

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Atoms and Molecules

15. Which of the following represents the correct relation between Avogadro's number (N_o), number of particles (N) and moles (n)?

(a) $n = N / N_o$ (b) $n = N_o / N$ (c) $n = N N_o$ (d) all are correct

16. The atomicity of $K_2Cr_2O_7$ is

(a) 9 **(b) 11** (c) 10 (d) 12

17. The formula for quicklime is

(a) $CaCl_2$ (b) $CaCO_3$ (c) $Ca(OH)_2$ **(d) CaO**

18. The symbol of cadmium is

(a) Ca (b) Cu (c) Cm **(d) Cd**

19. All noble gas molecules are

(a) Monoatomic (b) Diatomic (c) Triatomic (d) Both I and II

20. The valency of nitrogen in NH3 is

(a) 1 **(b) 3** (c) 4 (d) 5

21. The formula of ethanol is C_2H_5OH. What will be its molecular mass?

(a) 46 u (b) 34 u (c) 34 g **(d) 46 g**

22. Number of moles present in 28g of nitrogen atoms are

(a) 1 mole (b) 2.3 moles (c) 0.5 mole **(d) 2 moles**

23. The molecular mass of x is 106. x can be

(a) $CaCO_3$ (b) SO_3 **(c) Na_2CO_3** (d) NaCl

24. Which among the following is not a postulate of Dalton's atomic theory?

(a) Atoms cannot be created or destroyed

(b) Atoms of different elements have different sizes, masses and chemical properties

(c) Atoms of same elements can combine in only one ratio to produce more than any one compound

(d) Atoms are very tiny particles which cannot be further divided

25. Which of the following is a wrong Combination?

(a) $6.022 * 10^{23}$ molecules of oxygen = 32g of oxygen

(b) $6.022 * 10^{23}$ ions of sodium = 23g of sodium

(c) $6.022 * 10^{23}$atoms of C = 24g of carbon

(d) $6.022 * 10^{23}$ atoms of H = 1g of hydrogen atoms

26. Which of the following has largest number of particles?

(a) 8g of CH_4 (b) 4.4g of CO_2

(c) 34.2g of $C_{12}H_{22}O_{11}$ **(d)2g of H_2**

27. The number of molecules in 16.0g of oxygen is:

(a)6.02×10^{23} (b)6.02×10^{-23} (c)3.01×10^{-23} **(d)3.01×10^{23}**

28. The percentage of hydrogen in H_2O is:

(a) 8.88 **(b)11.12** (c)20.60 (d)80.0

29. Find the mass of oxygen contained in 1 kg of potassium nitrate KNO3.

(a)475.5g (b)485.5g **(c)475.2g** (d)485.2g

30. How many moles of electron weight one kilogram?

(a) 6.023×10^{23} (b) 6.023×10^{8} (c) 9.108×10^{54} **(d) 1.82×10^{6}**

31. 25.4g of iodine and 14.2g of chlorine are made to react completely to yield a mixture of ICl and ICl_3. Calculate the ratio of moles of ICl and ICl_3:

(a)1: 1 (b)2: 1 (c)3: 1 (d)1: 2

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